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NON-SOLVATED ALUMINUM HYDRIDE

Technical Report to the Office of Naval Research  
Contract ONR-494(04)

By

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ABSTRACT

The preparation of non-solvated aluminum hydride is described. The properties of the compound are described and results of X-ray diffraction studies are given.

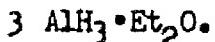
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A. Introduction

The classical method for preparing aluminum hydride is based on the reaction between ether solutions of aluminum chloride and lithium aluminum hydride (1).



This results in an ether solution of aluminum hydride which, within a short time, comes out of solution as a curdy, white solid which is presumed to be polymeric. The solid gives up some of the co-ordinated ether when treated under vacuum, yielding a material of the composition:



Further vacuum treatment and heating decompose the material rather than removing the remaining ether; pure, non-solvated aluminum hydride must be prepared by some other method. Stecher and Wiberg (2) prepared non-solvated aluminum hydride by treating aluminum alkyls in a hydrogen glow discharge. The product was not pure in that all possible aluminum alkyl hydrides were present and separation was carried out by preparing coordination compounds with amines. Also, the method was not practical for preparing useful quantities of pure aluminum hydride.

Several possible structures of aluminum hydride have been proposed. Each of these neglects the co-ordinated ether, however, which is present in the polymeric form. Wiberg (2) first proposed a linear structure in which the aluminum atoms are connected by single hydrogen bridges:



A later proposal by Wiberg (3) was essentially a simplified version of one proposed earlier by Lonquet-Higgins (4) in which hexagons of aluminum atoms lie in a plane, with two bridge hydrogens between aluminum atoms lying above and below the plane of the hexagons.

The present investigation was carried out as a portion of a more general study of aluminum hydride. The specific purpose of this study was to prepare non-solvated aluminum hydride and to attempt to determine its structure.

B. Experimental and Results

Two basic ideas were used in attempting to prepare non-etherated aluminum hydride; first, various solvents other than diethyl ether were used in preparing the aluminum hydride and second, aluminum hydride solutions in diethyl ether were run in to various solvents and the solvent mixtures removed by vacuum.

Both diisopropyl ether and dimethyl sulfide were tried as solvents in lieu of diethyl ether. In each case reaction was incomplete and in the case of dimethyl sulfide a mixture of products was obtained indicating that undesirable side reactions were occurring. These results gave no indication of success and this type of attack was abandoned.

A great deal of time and effort were devoted to studying the possibility of ridding aluminum hydride of ether by treatment with other inert solvents. At first, the study consisted of adding ether solutions of aluminum hydride to various solvents and removing the mixed solvents in vacuo in order to single out the most promising "ether remover". Table I lists the results obtained with various solvents under similar conditions

Table I

Precipitating Solvent	Volume Solvent Per Gram AlCl <sub>3</sub>	Remarks	%H	% Al
Benzene	19 ml.	Evacuated for a few hours at room temp.	1.76	14.70
Dioxane	19 ml.	"	1.89	20.42
Pentane	19 ml.	"	5.15	48.67
Tetrahydrofuran (no ppt.)	20 ml.	"	1.57	15.7
Carbon disulfide	20 ml.	"	Exploded	
		Theoretical	10.0	99.0

The results of these early experiments showed that pentane was the most promising of all the solvents tried. Accordingly, varying conditions were tried until it was found possible to obtain non-etherated aluminum hydride. It was found that the volume of inert solvent used was critical. At least, 100 ml. of pentane per gram of AlCl<sub>3</sub> (used in making the AlH<sub>3</sub>) was necessary.

A typical preparation is as follows:

The apparatus used in the preparation of the non-solvated hydride is shown in Figure 1. The reaction flask B was taken into a drybox and the desired amount of aluminum chloride weighed out carefully on a triple beam balance. The flask was tightly stoppered and then connected to the apparatus. A quantity of ether,

sufficient to provide approximately 10 ml. of ether per gram of  $\text{AlCl}_3$ , was added slowly through the top of the condenser A. A drying tube was added to the top of the condenser and the solution stirred until all of the  $\text{AlCl}_3$  had dissolved. A stoichiometric amount of  $\text{LiLiH}_3$  solution was added and after briefly stirring the  $\text{LiCl}$  was allowed to settle. Nitrogen pressure was applied at the top of the condenser to force the  $\text{AlH}_3$  solution through the fritted glass filter C into flask D containing pentane. The solution was run down a wire in a thin film into the pentane which was stirred by a magnetic stirrer. The  $\text{AlH}_3$  precipitated immediately on contact with the pentane. After addition of the  $\text{AlH}_3$  solution the reaction flask and filter were removed from the system. The solvent mixture was removed by vacuum and collected in the cold trap E. It was usually necessary to maintain the  $\text{AlH}_3$  under vacuum for at least 12 hours at room temperature in order to remove all the ether. Ether removal was facilitated by heating the flask to 50°C. and by repeatedly grinding the product with the magnetic stirrer. When the pressure in the system became less than 2  $\mu$ , the system was filled with nitrogen, flask D removed and stoppered and the flask removed to a nitrogen drybox where it could be handled safely. Freshly prepared samples were white, fluffy, pyrophoric powders with a low bulk density. The samples were fairly stable at room temperature. Stability seemed to be related to purity; the container with the most pure sample developed negligible hydrogen pressure over a period of weeks while some samples containing chloride turned jet black in a few days.

Analysis of the products was carried out using standard procedures. Hydrogen was determined by evolution with water and aluminum by the 8-hydroxyquinoline method. A test for chloride was made using silver nitrate; if more than a trace of chloride was present, it was determined by the Volhard method using eosin as an indicator. All samples gave a positive flame test for lithium but lithium was present in too small amounts to be determined by wet methods.

Densities of aluminum hydride samples were determined in a helium densitometer similar to one described by Schumb (5). The density as a function of ether content is shown in Figure 2. The density of the pure hydride was determined to be 1.718 g/ml, but this value was subject to some error due to the small weight of sample used in the measurement.

X-ray powder patterns were made of many samples using Cu-K $\alpha$  radiation in a General Electric XRD-4 apparatus. No patterns were obtained unless the samples were essentially non-etherated. In Table 2 are listed d-spacings obtained from several samples.

Table 2,

d-Spacings For Aluminum Hydride Samples

<u>Sample No. 113 94% Hydride</u>	<u>Sample No. 115 ≥ 96% Hydride</u>	<u>Sample A-16 87% Hydride</u>
5.34	5.34	4.48
4.57	4.57	4.30
4.35	4.33	4.14
3.67	3.75	3.32
3.50	3.46	3.23
3.22	3.23	2.67
3.04	3.06	2.46
2.88	2.88	2.43
2.69	2.76	2.24
2.40	2.47	2.15
2.31	2.32	2.10
2.26	2.28	2.05
2.08	2.08	1.98
1.99	1.99	1.92
1.76	1.76	1.78
1.71	1.71	1.74
1.66	1.61	1.68
1.53	1.57	1.62
1.44	1.53	1.52
1.34	1.44	1.43
1.23		1.25
		1.15

C. Discussion

It is worthwhile to speculate on the mechanism of polymerization of aluminum hydride. Zeil (6) determined force constants for Al-H bonds and Al-O bonds in the co-ordination compounds of aluminum hydride with tetrahydrofuran. For the Al-H bond  $k = 1.8 \times 10^5$  dynes/cm. and for the Al-O bond  $k = 4.0 \times 10^5$  dynes/cm. The fact that the Al-O bond is about twice as strong as the Al-H bonds makes it clear why the aluminum hydride etherate decomposes, giving off hydrogen, while still retaining ether. It also indicates that the freshly prepared aluminum hydride is not co-ordinated to ether since running it into an inert solvent would not break up the co-ordination compound. It seems likely that the "polymerization" which is observed in aluminum hydride solutions is connected with,

or may actually be, the formation of the etherate. Of course, this does not indicate the nature of the dissolved species other than it is not co-ordinated to the ether.

The action of the inert solvent or precipitating agent merely seems to precipitate the aluminum hydride before it co-ordinates with the ether. Once the hydride had been precipitated it remained non-etherated if it was precipitated as a non-etherate. In one experiment, a sample of the non-solvated hydride was treated with ether. After several hours the ether was removed by heating to 50°C. in vacuum. Upon re-analysis there was found to be no essential change in the material from the time that the non-solvated material was first made. This shows that once the aluminum hydride is non-solvated it will not co-ordinate with ether.

Although pentane was used most frequently as the precipitating agent, other agents were found to be suitable. Hexane and benzene were tried successfully although benzene took longer to remove after precipitation due to its higher boiling point. Presumably other liquids, which would not react with or act as a solvent for aluminum hydride, would also be suitable.

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